

Our Ref.: KB-038-US-X

TITLE OF THE INVENTION

Inventor
COOKING METHOD FOR PULP

BACKGROUND OF THE INVENTION

5 FIELD OF THE INVENTION

The present invention relates to a method for cooking a lignocellulose material, particularly to an effective cooking method for pulp, wherein a polysulfide cooking liquor and a quinone compound are used in combination.

10 DISCUSSION OF BACKGROUND

The principal method for producing chemical pulp which has heretofore been industrially employed, is an alkaline cooking method of a lignocellulose material such as wood chip, whereby a kraft method employing an alkaline cooking liquor comprising sodium hydroxide and sodium sulfide as the main components, has been used in many cases. Further, as one of cooking methods to improve the yield of pulp, a so-called polysulfide cooking method is widely known, wherein cooking is carried out by means of an alkaline cooking liquor containing polysulfides. According to this polysulfide cooking method, polysulfide ions oxidize and stabilize terminal aldehyde groups of cellulose and hemi-cellulose, to prevent a peeling reaction and to suppress a reaction for elution of cellulose and hemi-cellulose, whereby the yield of pulp will be improved. And, in general, the

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higher the concentration of the polysulfide sulfur in this polysulfide cooking liquor, the higher the cooking effects.

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The alkaline cooking liquor containing polysulfides, to be used in the above cooking method, is produced by a method of air oxidation in the presence of a catalyst (for example, JP-B-50-40395, JP-A-61-257238, JP-A-61-259754, JP-A-09-87987). In this method, when usual white liquor is employed, it is possible to obtain an alkaline cooking liquor having a polysulfide sulfur concentration of about 5 g/l (l represents liter, the same applies in this specification) at a reaction rate of about 60% and a selectivity of about 60%. However, during the formation of polysulfides, this method produces thiosulfate ions as a by-product which is ineffective for cooking, whereby it has been difficult to produce an alkaline cooking liquor containing highly concentrated polysulfide sulfur at a high selectivity.

On the other hand, as shown in e.g. JP-B-57-19239, JP-B-53-45404 and JP-A-52-37803, a quinone cooking method is also widely known, wherein cooking is carried out by adding a quinone-hydroquinone compound to an alkaline cooking liquor. The added quinone compound oxidizes and stabilizes the terminal aldehyde groups of cellulose and hemi-cellulose, thereby to prevent a peeling reaction and suppress an elution reaction of cellulose and hemi-cellulose. On the other hand, the quinone compound which

has become a hydroquinone type, will act on lignin to reduce and elute the lignin and to become a quinone type itself. Thus, the quinone-hydroquinone compound stabilizes cellulose and hemi-cellulose and accelerates delignification by the oxidation-reduction cycle of itself, whereby even when compared under such a condition that the Kappa number of pulp is the same, it brings about effects to improve the yield and at the same time to reduce the amount of active alkali required for cooking. Here, in this specification, the quinone-hydroquinone compound means both a quinone compound as an oxidation type quinone substance and a hydroquinone compound as a reduction type hydroquinone substance.

In the Journal of Japan Technical Association of Pulp and Paper Industry, Vol. 32, No. 12, p. 713-721 (1978), Nomura et al. disclose that in cooking for kraft pulp employing a cooking liquor comprising sodium hydroxide and sodium sulfide as the main components, which is commonly adopted as a cooking method for pulp, if a quinone compound is employed, of which the oxidation-reduction potential in the form present during the cooking, which potential is a value calculated as a standard oxidation-reduction potential (E_a) with a hydrogen ion activity of 1, is from 0.1 to 0.25V to the standard hydrogen electrode potential, it is possible to improve the yield, etc. of pulp, and they disclose that even within this potential range, a quinone compound such

as anthraquinone carboxylic acid or anthraquinone dicarboxylic acid having a potential higher than 9,10-anthraquinone ($E_a=0.154V$) is inferior in the effects, and a quinone compound such as hydroxyanthraquinone having a low potential has larger effects than 9,10-anthraquinone.

Further, as shown in e.g. JP-A-7-189153, a so-called polysulfide-quinone cooking method having the above-mentioned cooking method combined, is also widely known. By this cooking method, the above-described effects appear synergistically. Namely, as effects of the polysulfide-quinone cooking, improvement in the yield of pulp as compared with the same Kappa number and reduction in the amount of active alkali to be used as compared with the same amount of pulp production, can be accomplished over the cases where the respective techniques are separately employed.

However, no research or development has been made on what type of quinone compounds is effective for cooking and for improvement in the yield of pulp or in the required amount of the chemical solutions to be used, in the presence of polysulfides. In the present invention, a research and study have been made on a cooking method relating to such an aspect, and as a result, it has been found that further improvement in the yield of pulp, further reduction in the amount of the chemical solutions to be used, and solution of the problem relating to the load on the recovery boiler, can be accomplished, thus

arriving at the present invention.

SUMMARY OF THE INVENTION

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The present invention provides a cooking method for pulp, which comprises polysulfide cooking method pulping
5 a lignocellulose material with an alkaline cooking liquor containing polysulfides in the presence of a quinone-hydroquinone compound, wherein the oxidation-reduction potential of the quinone-hydroquinone compound in the form present during the cooking, which potential is a
10 value calculated as a standard oxidation-reduction potential (E_a) with a hydrogen ion activity of 1, is from 0.12 to 0.25V to the standard hydrogen electrode potential.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

15 In the present invention, it is important that in the cooking method for pulping a lignocellulose material with an alkaline cooking liquor containing polysulfides in the presence of a quinone-hydroquinone compound, the oxidation-reduction potential of the quinone-hydroquinone
20 compound in the form present during the cooking, which potential is a value calculated as a standard oxidation-reduction potential (E_a) with a hydrogen ion activity of 1, is made to be from 0.12 to 0.25V to the standard hydrogen electrode potential, According to the present
25 invention, as compared with a kraft cooking method or a cooking method having a kraft cooking combined with either polysulfides or a quinone-hydroquinone compound

alone, it is possible to obtain effects to improve the yield and effects to reduce the amount of active alkali to be contained in the alkaline cooking liquor, as compared with the same Kappa number of the obtained pulp.

5 In addition thereto, it is possible to obtain effects to increase the production as the cooking time can be shortened and to obtain a merit such that the cooking effects scarcely deteriorate even when the liquid to wood ratio is increased.

10 In the present invention, an alkaline cooking liquor containing polysulfides, is employed. By the oxidation action of the polysulfide sulfur contained in the polysulfide cooking liquor, it is possible to accelerate the stabilization of cellulose and hemi-cellulose and to
15 improve the yield of pulp. Here, a polysulfide ion is represented by the general formula S_x^{2-} and may simply be referred to as a polysulfide. The polysulfide sulfur is meant for sulfur having an oxidation number of 0 in sulfur atoms constituting polysulfide ions and sulfur of
20 (x-1) atoms in S_x^{2-} . Further, Na_2S -state sulfur generally refers to sulfur having oxidation number of -II in the polysulfide ions (sulfur of one atom per S_x^{2-}) and sulfide ions. Further, the active alkali is $NaOH+Na_2S$ calculated as a Na_2O concentration.

25 In the present invention, as the quinone-hydroquinone compound to be used in this polysulfide-quinone cooking method, one having a standard oxidation-

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reduction potential (Ea) in the form present during the cooking within a range of from 0.12 to 0.25V, is employed. It is more preferred to select one having a standard oxidation-reduction potential within a range of from 0.14 to 0.20V, whereby further improvement in the cooking effects can be obtained. Here, the standard oxidation-reduction potential is a potential represented by a value obtained by converting the oxidation-reduction potential in the form present during the cooking, into a standard oxidation-reduction potential (Ea) with a hydrogen ion activity of 1, against the standard hydrogen electrode potential.

As mentioned above, the Journal of Japan Technical Association of Pulp and Paper Industry, Vol. 32, No. 12, p. 713-721 (1978) discloses that in a kraft pulp cooking employing a cooking liquor comprising sodium hydroxide and sodium sulfide as the main components which is commonly adopted as a cooking method for pulp, if a quinone compound is employed, of which the oxidation-reduction potential in the form present during the cooking, which potential is a value calculated as the standard oxidation-reduction potential (Ea) with a hydrogen ion activity of 1, is from 0.1 to 0.25V to the standard hydrogen electrode potential, it is possible to improve the yield, etc. of pulp, and it further discloses that even within this potential range, a quinone such as anthraquinone carboxylic acid or anthraquinone

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dicarboxylic acid having a potential higher than 9,10-anthraquinone ($E_a=0.154V$), is inferior in the effects, and a quinone such as hydroxyanthraquinone having a low potential has larger effects than 9,10-anthraquinone.

5 However, no substantial research or study has been made on a combination of a quinone compound with the polysulfide cooking. In general, the effects of the quinone compound are such that as mentioned above, the quinone compound oxidizes and stabilizes the terminal
10 aldehyde groups of cellulose and hemi-cellulose, whereby the peeling reaction is prevented to suppress the reaction for elution of cellulose and hemi-cellulose. On the other hand, the quinone compound which has become a hydroquinone type acts on lignin to reduce and elute the
15 lignin and becomes a quinone type itself. Thus, the quinone-hydroquinone compound has effects to stabilize cellulose and hemi-cellulose and to accelerate delignification by the oxidation-reduction cycle of itself. If polysulfide ions are added thereto, the
20 polysulfide ions have effects to oxidize and stabilize the terminal aldehyde groups of cellulose and hemi-cellulose, whereby the quinone capable of effectively promoting delignification is believed to be more effective.

25 Namely, in a so-called polysulfide-quinone cooking method, a quinone-hydroquinone compound having a large reduction power is advantageous. It is easily assumed

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that oxidation and stabilization of cellulose and hemi-cellulose are thereby accelerated, and the range of the standard oxidation-reduction potential of the quinone compound to further improve the cooking effects will
5 shift to a range lower than from 0.1 to 0.25V.

However, as a result of experiments conducted by the present inventors on polysulfide cooking employing quinone-hydroquinone compounds having various standard oxidation-reduction potentials, it has been found, as
10 totally contrast to the above assumption, that if the standard oxidation-reduction potential is lower than 0.12V, no substantial cooking effects will appear. Namely, by many experiments, it has been made clear that if the standard oxidation-reduction potential of the
15 quinone-hydroquinone compound becomes lower than 0.12V, the effects to improve the yield of pulp and effects to reduce the amount of the active alkali to be used, tend to decrease, and if the standard oxidation-reduction potential exceeds 0.25V, the effects to improve the yield
20 of pulp and the effects to reduce the amount of the active alkali to be used, tend to decrease. The value is more preferably within a range of from 0.14V to 0.20V. The present invention is applicable not only to a usual kraft method but also all cooking methods for pulp,
25 including a modified kraft method (MCC method) and a Lo-Solids (registered trademark) method.

In the present invention, the quinone-hydroquinone

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compound, of which the oxidation-reduction potential in the form present during the cooking, which potential is a value calculated as a standard oxidation-reduction potential (E_a) with a hydrogen ion activity of 1, is from 0.12 to 0.25V to the standard hydrogen electrode potential, may specifically be, for example, an alkyl anthraquinone such as 1-ethyl-9,10-anthraquinone ($E_a=0.140V$), 9,10-anthraquinone ($E_a=0.154V$) or 2-methyl-9,10-anthraquinone ($E_a=0.150V$), a quinone compound such as 1-hydroxy-9,10-anthraquinone ($E_a=0.140V$), 2-(9,10-anthraquinoyl)-1-ethanesulfonic acid ($E_a=0.162V$), 9,10-anthraquinone-2-sulfonic acid ($E_a=0.187V$), 9,10-anthraquinone-2-carboxylic acid ($E_a=0.213V$), 9,10-anthraquinone-2,7-disulfonic acid ($E_a=0.228V$), benz(α)anthracene-7, 12-dione ($E_a=0.228V$), 1,4,4a,9a-tetrahydro-9,10-anthraquinone ($E_a=0.154V$) or 1,4-dihydro-9,10-anthraquinone ($E_a=0.154V$), and hydroquinone compounds as reduction products thereof.

These standard oxidation-reduction potentials E_a are taken from or in accordance with "Dai Yukikagaku Bekkan 2, Yukikagaku Josu Binran", published by Asakura Shoten, p. 670-680 (1963). The oxidation-reduction potentials of these quinone compounds can be measured by e.g. a usual method employing a cyclic voltammetry, but taking into an error by a measuring apparatus or a measuring person, it is necessary to calculate the measured value by using as a standard, an anthraquinone of which the potential is

known, such as 9,10-anthraquinone.

When such a quinone compound is added, it may be an oxidation type quinone substance or a reduction type hydroquinone substance. Irrespective of the state at the time of the addition, it is only required that the quinone-hydroquinone compound in the form present at the time of the cooking is within the above-mentioned potential range. For example, 1,4,4a,9a-tetrahydro-9,10-anthraquinone is present in the form of a disodium salt of 1,4-dihydro-9,10-dihydroxyanthracene in an alkaline cooking liquor. This will be readily oxidized at the initial stage of the cooking to 1,4-dihydro-9,10-anthraquinone, which is further readily transferred to 9,10-anthrahydroquinone, and during the cooking, it is acting in the form of 9,10-anthraquinone and 9,10-anthrahydroquinone. The same applies to 1,4-dihydro-9,10-anthraquinone.

In the present invention, the higher the polysulfide sulfur contained in the polysulfide cooking liquor, the higher the cooking effects. Accordingly, it is preferred to prepare the liquor so that the concentration of the polysulfide sulfur contained in the polysulfide cooking liquor becomes to be at least 6 g/l, more preferably at least 8 g/l.

In the present invention, as a method for producing the polysulfide cooking liquor, it is possible to employ a conventional air oxidation method. However, when a

polysulfide cooking liquor containing polysulfide sulfur is produced by the air oxidation method, formation of sodium thiosulfate as a by-product tends to be large, such being disadvantageous. Accordingly, it is preferred to employ a method of electrically oxidizing an alkaline solution containing sulfide ions, i.e. to form the cooking liquor by electrolysis. By such a method, it is possible to produce a polysulfide cooking liquor having a high concentration of a level of at least 8 g/l at a high selectivity. As such an electrolytic method, an electrolytic method of e.g. PCT/JP97/01456, JP-A-10-166374, JP-A-11-51016 or JP-A-11-51033 which has previously been developed by the present inventors, may be employed.

As the electrolytic cell to be used for the electrolytic method, a two compartment type electrolytic cell comprising one anode compartment and one cathode compartment, is required, or one having three or more compartments combined, may be employed. A plurality of electrolytic cells may be arranged to have a monopolar structure or a bipolar structure. To the anode compartment, an alkaline solution containing sulfide ions is introduced, and some sulfide ions are oxidized to form polysulfide ions. And, alkali metal ions will be transferred through a diaphragm to the cathode compartment.

On the other hand, into the cathode compartment,

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water or a solution comprising water and an alkali metal hydroxide, is introduced, so that the reaction for forming hydrogen gas from water, is preferably utilized. As a result, from the formed hydroxide ions and alkali metal ions transferred from the anode compartment, an alkali metal hydroxide will be formed. The concentration of the alkali metal hydroxide in the cathode compartment is, for example, from 1 to 15 mol/l, preferably from 2 to 5 mol/l. The anode disposed in the anode compartment of the electrolytic cell is preferably such that the entirety of the anode or at least the surface portion thereof, is made of a material excellent in alkali resistance. For example, nickel, titanium, carbon or platinum has practically adequate durability in the production of polysulfides. With respect to the structure of the anode, it is preferred to use a porous anode which is porous and has a three dimensional network structure. Specifically, a foam or an aggregate of fibers may, for example, be mentioned. Such a porous anode has a large surface area, whereby the desired electrolytic reaction takes place over the entire surface of the electrode surface, and formation of a by-product can be suppressed.

The surface area of the anode to be used for the electrolytic method, is preferably from 2 to 100 m²/m² in the case where the anode is a foam and from 30 to 5,000 m²/m² in a case where the anode is an aggregate of

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fibers, per unit area of the diaphragm partitioning the anode compartment and the cathode compartment. More preferably, it is from 5 to 50 m²/m² and 70 to 1,000 m²/m², respectively. If the surface area is too small, the current density at the anode surface tends to be large, whereby not only a by-product such as thiosulfate ions is likely to form, but also dissolution of the anode is likely to take place, such being undesirable. If the surface area is made to be too large, there will be a problem from the viewpoint of electrolytic operation such that the pressure loss of the liquid tends to be large, such being undesirable.

The average pore diameter of the network of the foam anode to be used for the electrolytic method is preferably from 0.1 to 5 mm. If the average pore diameter of the network is larger than 5 mm, the surface area of the anode can hardly be made large, whereby the current density at the anode surface tends to be large, and a by-product such as thiosulfate ions is likely to form, such being undesirable. If the average pore diameter of the network is smaller than 0.1 mm, there will be a problem from the viewpoint of electrolytic operation such that the pressure loss of the liquid tends to be large, such being undesirable. The average pore diameter of the network of the anode is more preferably from 0.2 to 2 mm.

With respect to the porous anode to be used for the

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electrolytic method, the diameter of the net constituting the network is preferably from 0.01 to 2 mm in the case of a foam and from 1 to 300 μm in the case of an aggregate of fibers. If the diameter is lower than the respective ranges, the production is very difficult and costly, and besides, handing will be difficult, such being undesirable. If the diameter exceeds the respective ranges, it is difficult to obtain an anode having a large surface area, whereby the current density at the anode surface will be large, and a by-product such as thiosulfate ions is likely to form, such being undesirable. Particularly preferably, the diameter is from 0.02 to 1 mm and from 5 to 50 μm , respectively.

The anode in the electrolytic cell may be disposed fully in the anode compartment so that it is in contact with the diaphragm. Otherwise, it may be disposed so that there will be a some space between the anode and the diaphragm. It is required that the liquid to be treated, flows in the anode, and accordingly, it is preferred that the anode has a sufficient porosity. In any case, the porosity of the anode is preferably from 90 to 99% in the case of a foam and from 70 to 99% in the case of an aggregate of fibers. If the porosity is too low, the pressure loss increases, such being undesirable. If the porosity exceeds 99%, it tends to be difficult to increase the surface area of the anode, such being undesirable. The porosity is more preferably from 90 to

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98% and from 80 to 95%, respectively.

With respect to the cathode to be used for the electrolytic method, the material is preferably an alkali resistant material, and nickel, Raney Nickel, nickel sulfide, steel or stainless steel may, for example, be employed. The shape may be a flat plate or meshed shape, and one or more may be employed in a multi-layer structure. A three dimensional electrode having a linear electrode combined, may also be employed.

As the diaphragm partitioning the anode compartment and the cathode compartment, to be used in the electrolytic method, it is preferred to employ a cation exchange membrane. The cation exchange membrane introduces cations from the anode compartment to the cathode compartment but prevents transfer of sulfide ions and polysulfide ions. As such a cation exchange membrane, a polymer membrane having cation exchange groups such as sulfonic groups or carboxylic groups introduced to a polymer of a hydrocarbon type or a fluorine type, is preferred. Further, a bipolar membrane or an anion exchange membrane may also be used if there is no problem with respect to the alkali resistance, etc.

In the electrolytic method, the operation is preferably carried out at a current density of from 0.5 to 20 kA/m² at the diaphragm surface. If the current density is less than 0.5 kA/m², an unnecessarily large electrolytic installation will be required, such being

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undesirable. If the current density at the diaphragm surface exceeds 20 kA/m^2 , by-products such as thiosulfate, sulfuric acid and oxygen, may increase, such being undesirable. The current density at the diaphragm surface is more preferably from 2 to 15 kA/m^2 . In the present electrolytic method, an anode having a large surface area to the area of the diaphragm, is employed, whereby operation can be carried out within a small range of the current density at the anode surface.

10 In the electrolytic method, the average superficial velocity in the anode compartment is preferably from 1 to 30 cm/sec. in the case of a foam and from 0.1 to 30 cm/sec. in the case of an aggregate of fibers. If the average superficial velocity is too small, the anode solution in the anode compartment will not be adequately stirred, and in some cases, precipitates are likely to deposit on the diaphragm facing the anode compartment, whereby the cell voltage is likely to increase as the time passes. Further, if it is larger than 30 cm/sec., the pressure loss will increase, such being undesirable. 20 The flow rate of the cathode solution is not particularly limited, but is determined by the degree of buoyancy of the generated gas. The temperature of the anode compartment is preferably from 70 to 110°C . If the temperature of the anode compartment is lower than 70°C , not only the cell voltage becomes high, but also dissolution of the anode or formation of by-products are 25

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likely to result, such being undesirable. The upper limit of the temperature is practically limited by the material of the diaphragm or the electrolytic cell. The solution containing sulfide ions to be introduced into the anode compartment is usually treated by one path or by recycling.

In the present invention, as a raw material for an alkaline cooking liquor containing polysulfides to be produced by the electrolytic method, it is preferred to employ white liquor or green liquor which is used at a pulp mill. In the case of white liquor currently employed for kraft pulp cooking, the composition of the white liquor usually contains from 2 to 6 mol/l of alkali metal ions, and at least 90% thereof is sodium ions, the rest being substantially potassium ions. Further, the anions include hydroxide ions, sulfide ions and carbonate ions as the main components, and the sulfide ion concentration is usually from 0.5 to 0.8 mol/l. Further, it contains sulfate ions, thiosulfate ions, chlorine ions and sulfite ions. Further, it contains trace amount components such as calcium, silicon, aluminum, phosphorus, magnesium, copper, manganese and iron. The composition of green liquor is basically the same as white liquor. However, while the white liquor contains sodium sulfide and sodium hydroxide as the main components, the green liquor contains sodium sulfide and sodium carbonate as the main components. In the

electrolytic method, a part of sulfide ions in such white liquor or green liquor is oxidized in the anode compartment to form polysulfide ions, which will be supplied to the cooking step.

5 In the present invention, the Na_2S -state sulfur concentration in the alkaline cooking liquor containing polysulfides is preferably at least 10 g/l as calculated as Na_2O . If this concentration is less than 10 g/l, the highly concentrated polysulfide sulfur of at least 8 g/l
10 tends to be unstable, and the Kappa number of the pulp obtained by cooking tends to increase, and the yield of pulp is likely to deteriorate.

In the present invention, the quinone-hydroquinone compound is preferably added to the alkaline-cooking
15 liquor so that it will be from 0.01 to 1.5 wt% based on the bone-dry chip. More preferably it is from 0.02 to 0.06 wt%. If the addition of the quinone compound is less than 0.01 wt%, the amount is too small, whereby the Kappa number of the pulp after cooking will not be
20 reduced, and the relation between the Kappa number and the yield of pulp will not be improved. Further, even if the quinone compound is added beyond 1.5 wt%, no further reduction of the Kappa number of pulp after cooking or no further improvement of the relation between the Kappa
25 number and the yield of pulp can be observed.

In the present invention, with respect to the timing for the addition of the quinone compound, a method of

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adding it all at once before cooking or during cooking,
or a method of adding it stepwisely in a divided fashion,
is effective. However, it is preferred to add it so that
the alkaline cooking liquor containing the quinone
5 compound will sufficiently penetrate into the chip.

Further, in the present invention, the liquid to
wood ratio during the cooking is preferably adjusted to
be from 1.5 to 5.0 l/kg based on bone-dry chip.
Particularly when soft wood chip is employed as the
10 lignocellulose material, the liquid to wood ratio is more
preferably from 1.5 to 3.5 l/kg, and when hard wood chip
is employed, it is more preferably from 2.5 to 5.0 l/kg.
If the liquid to wood ratio is less than 1.5 l/kg, the
alkaline cooking liquor may not sufficiently penetrate
15 into the chip, whereby the cooking effects are likely to
deteriorate, such being undesirable. If the liquid to
wood ratio exceeds 5.0 l/kg, the effects to reduce the
amount of the chemical solutions to be used tend to be
low, such being undesirable.

20 Here, the liquid to wood ratio means the amount of
the liquid based on the weight of bone-dry chip in the
case of a batch system digester, and it means the ratio
of the amount by volume of the liquid flowing into the
digester to the amount by weight of bone-dry chip flowing
25 into the digester, per unit time, in the case of a
continuous system digester.

As the lignocellulose material to be used in the

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present invention, soft wood or hard wood chip may be used, and any type of tree may be employed. For example, the soft wood may, for example, be Cryptomeria (Japan cedar), Picea (Yezo spruce, Hondo spruce, Norway spruce, 5 Sitka spruce, etc.), Pinus (Monterey pine, Japanese red pine, Japanese black pine, etc.), Thuja (Western red cedar, Japanese arbovitae, etc.) or Tsuga (Japanese hemlock, Western hemlock, etc.), and the hard wood may, for example, be Eucalyptus (eucalyptus trees), Fagus 10 (beech trees), Quercus (oak, white oak, etc.) or Acacia (acacia trees).

EXAMPLES

Now, the present invention will be described in detail with reference to Examples, but the present 15 invention is by no means restricted by such specific Examples. Test methods were as follows.

Test methods

With respect to the yield of pulp of obtained unbleached pulp, the yield of cleaned pulp having lump 20 removed, was measured. The Kappa number of the unbleached pulp was determined in accordance with TAPPI test method T236hm-85. The quantitative analyses of sodium thiosulfate, Na_2S -state sulfur and polysulfide sulfur concentration calculated as sulfur, in the 25 alkaline cooking liquor, were carried out in accordance with the method disclosed in JP-A-7-92148.

EXAMPLE 1

(1) Preparation of a polysulfide cooking liquor

A two compartment electrolytic cell was assembled, which comprised a nickel plate as an anode current collector, a nickel foam as an anode (100 mm × 20 mm × 4 mm, average pore diameter of network: 0.51 mm, surface area of the anode per volume of the anode compartment: 5,600 m²/m³, surface area to the diaphragm area: 28 m²/m²), an iron expansion metal as a cathode and a fluororesin type cation exchange membrane as a diaphragm. The anode compartment had a height of 100 mm, a width of 20 mm and a thickness of 4 mm, and the cathode compartment had a height of 100 mm, a width of 20 mm and a thickness of 5 mm. The effective area of the diaphragm was 20 cm². Using model white liquor, circulation electrolysis was carried out at an anode solution linear velocity of 4 cm/sec. at a current density of 6 kA/m² at an electrolysis temperature of 90°C, whereby a polysulfide cooking liquor having the following composition was obtained at a selectivity of 97%.

Sodium hydroxide: 85.5 g/l (calculated as Na₂O)

Na₂S-state sulfur: 12.0 g/l (calculated as Na₂O)

Sodium carbonate: 15 g/l (calculated as Na₂O)

Sodium thiosulfate: 0.5 g/l (calculated as Na₂O)

Polysulfide sulfur: 9.0 g/l (calculated as sulfur)

(2) Cooking test

As a lignocellulose material, 25g of Japanese red

pine chip (25g by bone-dry weight) was used, and the above-mentioned polysulfide cooking liquor was added thereto so that the addition of active alkali would be 16 and 18 wt% (based on the bone-dry chip; calculated as Na₂O). The liquid to wood ratio was adjusted to be 2.7 l/kg based on the bone-dry chip, including the moisture brought in by the chip and distilled water added as the case requires. Cooking was carried out under such conditions that 9,10-anthraquinone (Ea=0.154V) as a quinone compound was added to the polysulfide cooking liquor so that it would be 0.05 wt% based on the bone-dry chip, the temperature was raised from 109°C to 170°C for 60 minutes, and the maximum temperature was maintained for 73 minutes. The results of the cooking are shown in Table 1. As compared with Comparative Examples 1 and 2, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 2

Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, tetrahydroanthraquinone (disodium 1,4-dihydro-9,10-dihydroxyanthracene, SAQ, trade name, manufactured by Kawasaki Kasei Chemicals Ltd.) (Ea=0.154V) was added so that it would be the same molar amount as in Example 1. The results of the cooking are shown in Table 1. Like in Example 1, as compared with Comparative Examples 1 and 2,

the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 3

5 Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, 2-methyl-9,10-anthraquinone ($E_a=0.154V$) was added so that it would be the same molar amount as in Example 1. The results of the cooking are shown in Table 1. Like in Example 1, as
10 compared with Comparative Examples 1 and 2, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 4

15 Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, sodium 9,10-anthraquinone-2-sulfonate ($E_a=0.187V$) was added so that it would be the same molar amount as in Example 1. The results of the cooking are shown in Table 1. Like in Example 1, as compared with Comparative Examples 1 and 2,
20 the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 5

25 Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, 1-hydroxy-9,10-anthraquinone ($E_a=0.125V$) was added so that it would be the same molar amount as in Example 1. The results of

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the cooking are shown in Table 1. Like in Example 1, as compared with Comparative Examples 1 and 2, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

5 EXAMPLE 6

Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, disodium 9,10-anthraquinone-2,7-disulfonate ($E_a=0.228V$) was added so that it would be the same molar amount as in Example 1. The results of the cooking are shown in Table 1. Like in Example 1, as compared with the Comparative Examples 1 and 2, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

15 COMPARATIVE EXAMPLE 1

Cooking was carried out in the same manner as in Example 1 except that the quinone compound or the like was not added. The results of the cooking are shown in Table 1.

20 COMPARATIVE EXAMPLE 2

Cooking was carried out in the same manner as in Example 1 except that as a quinone compound, 1,2-dihydroxy-9,10-anthraquinone ($E_a=0.107V$) was added so that it would be the same molar amount as in Example 1. The results of the cooking are shown in Table 1.

EXAMPLE 7

Cooking was carried out under the following

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conditions. Cooking was carried out in the same manner as in Example 1 except that as a lignocellulose material, 35g of beech chip (as bone-dry) was used. As a quinone compound, 9,10-anthraquinone ($E_a=0.154V$) was added to the polysulfide cooking liquor before raising the temperature in an amount of 0.05 wt% based on the bone-dry chip. The results of the cooking are shown in Table 2. As compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 8

Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, tetrahydroanthraquinone (disodium 1,4-dihydro-9,10-dihydroxyanthracene, SAQ, trade name, manufactured by Kawasaki Kasei Chemicals Ltd.) ($E_a=0.154V$) was added so that it would be the same molar amount as in Example 7. The results of the cooking are shown in Table 2. Like in Example 7, as compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 9

Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, 2-methyl-9,10-anthraquinone ($E_a=0.154V$) was added so that it would be the same molar amount as in Example 7. The results of

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the cooking are shown in Table 2. Like in Example 7, as compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

5 EXAMPLE 10

Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, 9,10-anthraquinone-2-sulfonic acid ($E_a=0.187V$) was added so that it would be the same molar amount as in Example 7.

10 The results of the cooking are shown in Table 2. Like in Example 7, as compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

15 EXAMPLE 11

Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, 1-hydroxy-9,10-anthraquinone ($E_a=0.125V$) was added so that it would be the same molar amount as in Example 7. The results of

20 the cooking are shown in Table 2. Like in Example 7, as compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

EXAMPLE 12

25 Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, disodium 9,10-anthraquinone-2,7-disulfonate ($E_a=0.228V$) was added

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so that it would be the same molar amount as in Example 7. The results of the cooking are shown in Table 2. Like in Example 7, as compared with Comparative Examples 3 and 4, the Kappa number at the same active alkali addition decreased, and the yield of pulp at the same Kappa number increased.

COMPARATIVE EXAMPLE 3

Cooking was carried out in the same manner as in Example 7 except that the quinone compound or the like was not added. The results of the cooking are shown in Table 2.

COMPARATIVE EXAMPLE 4

Cooking was carried out in the same manner as in Example 7 except that as a quinone compound, 1,2-dihydroxy-9,10-anthraquinone ($E_a=0.107V$) was added so that it would be the same molar amount as in Example 7. The results of the cooking are shown in Table 2.

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Table 1: Cooking test using soft wood (Japanese red pine) chip

Ex. No. and Comp. Ex. No.	Quinone compound	Standard oxidation- reduction potential (E0/V)	Active alkali addition=16%		Active alkali addition=18%		Yield of pulp (%) when the Kappa number=22
			Kappa number	Yield of pulp (%)	Kappa number	Yield of pulp (%)	
Ex. 1	9,10-anthraquinone	0.154	27.2	49.2	20.8	48.3	48.8
Ex. 2	Tetrahydroanthraquinone	0.154	25.4	49.2	20.3	48.4	49.0
Ex. 3	2-methyl-9,10-anthraquinone	0.154	25.0	48.5	20.8	47.8	48.4
Ex. 4	Sodium 9,10-anthraquinone-2-sulfonate	0.187	29.5	49.0	23.0	48.1	48.2
Ex. 5	1-hydroxy-9,10-anthraquinone	0.125	30.4	49.0	23.9	48.1	48.1
Ex. 6	Disodium 9,10-anthraquinone-2,7-disulfonate	0.228	30.7	49.1	24.2	48.1	48.1
Comp. Ex. 1	Not added	-	32.3	48.6	25.0	47.2	46.9
Comp. Ex. 2	1,2-dihydroxy-9,10-anthraquinone	0.107	31.3	48.9	24.3	47.3	47.2

The active alkali addition is represented by wt% based on bone-dry chip, as calculated as Na_2O .

Table 2: Cooking test using hard wood (beech) chip

Ex. No. and Comp. Ex. No.	Quinone compound	Standard oxidation- reduction potential (E0/V)	Active alkali addition=16%		Active alkali addition=18%		Yield of pulp (%) when the Kappa number=18
			Kappa number	Yield of pulp (%)	Kappa number	Yield of pulp (%)	
Ex. 7	9,10-anthraquinone	0.154	18.7	57.7	14.2	55.9	57.6
Ex. 8	Tetrahydroanthraquinone	0.154	17.9	57.7	13.6	56.1	57.8
Ex. 9	2-methyl-9,10-anthraquinone	0.154	20.2	57.8	13.6	55.7	57.4
Ex. 10	Sodium 9,10-anthraquinone-2-sulfonate	0.187	23.4	58.3	15.3	55.8	57.1
Ex. 11	1-hydroxy-9,10-anthraquinone	0.125	24.5	57.9	16.6	55.9	56.5
Ex. 12	Disodium 9,10-anthraquinone-2,7-disulfonate	0.228	27.0	58.1	17.3	55.9	56.1
Comp. Ex. 3	Not added	-	29.3	57.5	18.1	55.3	55.3
Comp. Ex. 4	1,2-dihydroxy-9,10-anthraquinone	0.107	28.7	57.8	17.5	55.1	55.5

The active alkali addition is represented by wt% based on bone-dry chip, as calculated as Na₂O.

INDUSTRIAL APPLICABILITY

According to the present invention, by pulping by means of an alkaline cooking liquor containing polysulfides, in the presence of a quinone-hydroquinone compound having a standard oxidation-reduction potential within a certain specific range, it is possible to further improve the yield of pulp and further improve the relation between the Kappa number and the yield of pulp. Namely, not only excellent effects are obtainable to reduce the Kappa number at the same active alkali addition and to improve the yield of pulp at the same Kappa number, but also effects to reduce the amount of chemical solutions to be used and effects to reduce the load on the recovery boiler, can be accomplished.

The entire disclosure of Japanese Patent Application No. 11-168948 filed on June 15, 1999 including specification, claims and summary are incorporated herein by reference in its entirety.

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